

3.

- 1.
- 2.
- 3.
- 4.
- 5.

1. (0,101080)

	, %	, / 3
	78.1	$9.76 \cdot 10^5$
	20.9	$2.98 \cdot 10^5$
	0.93	$1.66 \cdot 10^4$
2	0.03	$5.89 \cdot 10^2$
	$10^{-3} - 10^{-6}$	20 - 0,5
N ₂ O	$5 \cdot 10^{-5}$	0.98
	$5 \cdot 10^{-5}$	0.045
	$2 \cdot 10^{-6}$	0.042

_____ : S₂, CH₄, CO, H₂S, HF, CO₂,
().

- _____ :
- NO_x, CO₂, SO₂, NH₃, H₂S, Cl₂, Br₂, HHal, _____, AsH₃, PH₃;
 - _____ ;
 - _____ ;
 - _____ (_____, _____, _____, _____);
 - _____.

- _____ (NH₃, _____);
 - _____ (_____, 1- _____, _____)
 - _____ ;
 - _____.
- _____ :
-r < $2 \cdot 10^{-5}$; _____ - $2 \cdot 10^{-5} - 1 \cdot 10^{-4}$; _____ > $1 \cdot 10^{-4}$.

2.

1979 _____ « _____ » --
5 _____ :
1) _____ ; 2) SO₂; 3) _____ ; 4) _____ ; 5) NO, NO₂, N₂O₄.

_____ 4 _____ (_____) - / 3):

1. _____ :
0.16 (_____ 0.03); _____ (VI) _____ CrO₃ 0.0015; _____ 0.07;
() **0.1** /100³ (_____); SeO₂ 0.1 / 3; Pb, Bi, Hg, Tl, Ni.

(1/2, . .)

5.

<250 °C:

1	(, FID)			10 / ()	10 ⁷
2	(, ID): + RbBr CsBr		N- P-	1 / (N) 5 / ()	10 ⁴
3	(, ID)			-	10 ⁷
4	()		1!	0,2 / (l)	10 ⁴
5	()		Cl HCl, S SO ₂ , N NH ₃ ;	1 / (Cl) 5 / (S)	10 ⁶ 10 ⁴
6	()			1	10 ⁵
7	(, FPD)		S — 394 , — 526		10 ⁴
8	()		C, H, O, N, S, Hal,	0.2 / 50 /	10 ⁴
9					10 ³
10	()			1 / - 1 /	10 ⁵

- 1.
- 2.
- 3.

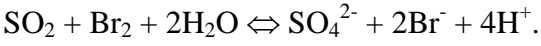
2.
 $Q = n_e \cdot F \cdot n(A)$, Q -
 n(A) -

$$i = n_e \cdot F \frac{n(A)}{t}, \quad \sim ();$$

$$\frac{n(A)}{t} = \frac{1}{M \cdot 10^3 (/)} \cdot (\frac{-}{3}) \cdot V (\frac{-}{3})$$

2	

:
 (Titrilog)
 :

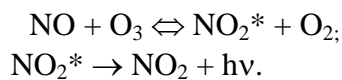


$$E = E^\ominus + \frac{2.303RT}{2F} \lg \frac{[Br_2]}{[Br^-]^2}$$

	FeSO ₄
NO ₂	KI
SO ₂	C NaOH
H ₂ S	Ag- 150°
NO	NO ₂ , NO ₂

- ;
- ;
- (NO → NO₂);
- (SO₂, NO₂, NO₃);
- NaOH (CO₂);
- Ag - (→ CO₂);
- (2);
-

2.
 :
 : ; SO₂, NO, H₂S, HF, Cl₂, CS₂,
 (SO₂, NO, Cl₂)
 (0,1 / 3)



H ₂ S	KIO ₃ , H ₂ SO ₄ Pb(Ac) ₂ , BaCl ₂ (VI)

3.

- 1.
- 2.
3. ().
4. -

— 1962 .
 n- : ZnO, SnO₂, Fe₂O₃, TiO₂, V₂O₇, MnO₂, CdO,
 p- : CoO, Cu₂O, NiO, Cr₂O₃.
 ZnO₂ CaO;
 Pt,

$$(O_2) \text{Pt} | \text{ZrO}_2, \text{CaO} | \text{Pt} (O_2)$$

$$P_1 > P_2,$$

$$E = \frac{2,303RT}{4F} \lg \frac{P_2}{P_1}, E \sim \lg P(O_2).$$

200

$\Delta F = -2,3 \cdot 10^6 F \frac{m}{S}$

$\Delta F -$; F - ; m - ; S -
 (HgO Hg),
 $\text{CO} \uparrow + \text{HgO} \downarrow \rightleftharpoons \text{CO}_2 \uparrow + \text{Hg} \uparrow.$
 SO₂:
 $\text{R}_3\text{N} + \text{SO}_2 + \text{H}_2\text{O} \rightleftharpoons \text{R}_3\text{NH}^+ \cdot \text{HSO}_3^-$
 - 550 () - 1000.